3.012 Fund of Mat Sci: Bonding – Lecture 1 bis

WAVE MECHANICS

Photo courtesy of Malene Thyssen, www.mtfoto.dk/malene/

Last Time

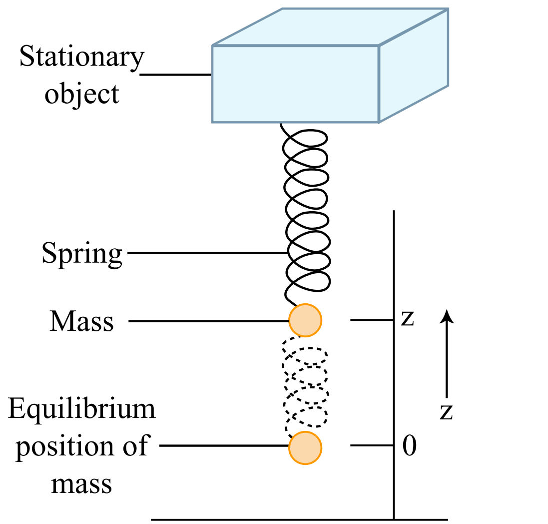
1. Players: particles (protons and neutrons in the nuclei, electrons) and electromagnetic fields (photons)
2. Forces: electromagnetic
3. Dynamics: Newton (macroscopic), Maxwell (fields), Schrödinger (microscopic)
4. De Broglie relation $\lambda \cdot p = h$

Homework for Mon 12

- You know by now: 12.5, 13.2, 13.3
- Study: 13.4 and 13.5

- Notes on harmonic oscillator -- Section 14.1in Mortimer, R. G. *Physical Chemistry*. 2nd ed. San Diego, CA: Elsevier, 2000.

Harmonic Oscillator (I)



A mass on a spring. This system can be represented by a harmonic oscillator.

Figure by MIT OCW.

Harmonic Oscillator (II)

Harmonic Oscillator (III)

Graph of the behavior of a harmonic oscillator removed for copyright reasons.
See Mortimer, R. G. *Physical Chemistry*. 2nd ed. San Diego, CA: Elsevier, 2000, p. 495, Figure 14.2.

The total energy of the system

- Kinetic energy K

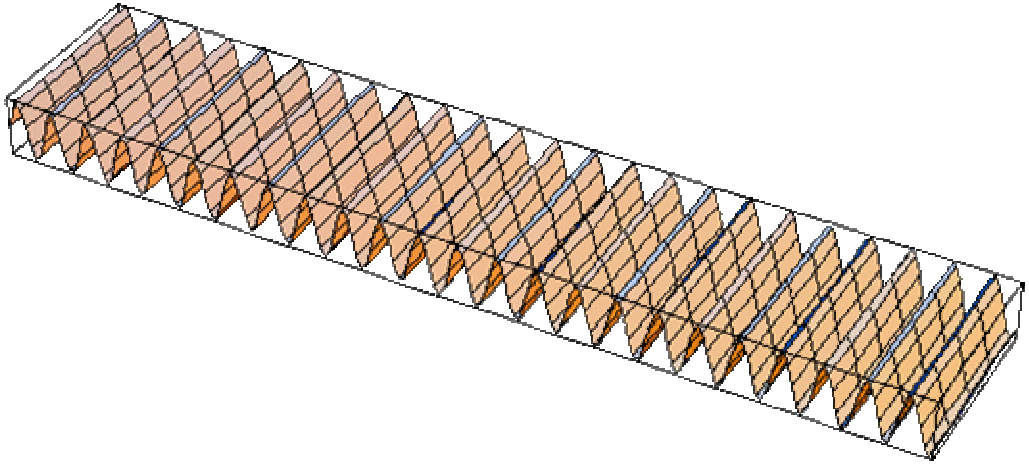
- Potential energy V

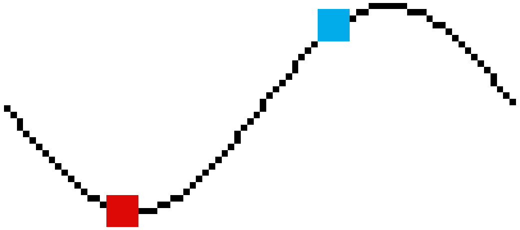
Polar Representation

Diagram of the Argand plane removed for copyright reasons.
See Mortimer, R. G. *Physical Chemistry*. 2nd ed. San Diego, CA: Elsevier, 2000, p. 1011, figure B.6.

A Traveling "Plane" Wave

$$\Psi(\vec{r},t) = A\exp[i(\vec{k}\cdot\vec{r}-\omega t)]$$





Principle of Linear Superposition

Photo courtesy of Spiralz.

Wave-particle Duality

- *Waves have particle-like properties:*
  - *Photoelectric effect: quanta (photons) are exchanged discretely*
  - *Energy spectrum of an incandescent body looks like a gas of very hot particles*

Diagrams of the photoelectric effect and of a P-N solar cell, removed for copyright reasons.

Wave-particle Duality

- Particles have wave-like properties:
  - Electrons in an atom act like standing waves (harmonics) in an organ pipe
  - Electrons beams can be diffracted, and we can see the fringes (Davisson and Germer, at Bell Labs in 1926…)

When is a particle like a wave ?

Wavelength • momentum = Planck

$\updownarrow$

Image of the double-slit experiment removed for copyright reasons.

See the simulation at http://www.kfunigraz.ac.at/imawww/vqm/movies.html:

"Samples from *Visual Quantum Mechanics*": "Double-slit Experiment."

$$\lambda \bullet p = h$$

$$(h = 6.626 \times 10^{-34} \text{ J s} = 2\pi \text{ a.u.})$$

See animation at http://www.kfunigraz.ac.at/imawww/vqm/movies.html
Select "Samples from *Visual Quantum Mechanics*" > "Double-slit experiment"

Time-dependent Schrödinger's equation

(Newton's 2nd law for quantum objects)

$$-\frac{\hbar^2}{2m}\nabla^2\Psi(\vec{r},t)+V(\vec{r},t)\Psi(\vec{r},t)=i\hbar\frac{\partial\Psi(\vec{r},t)}{\partial t}$$

1925-onwards: E. Schrödinger (wave equation), W. Heisenberg (matrix formulation), P.A.M. Dirac (relativistic)

Plane waves as free particles

Our free particle $\Psi(\vec{r},t) = A\exp[i(\vec{k}\cdot\vec{r} - \omega t)]$ satisfies the wave equation:

$$-\frac{\hbar^2}{2m}\nabla^2\Psi(\vec{r},t) = i\hbar\frac{\partial\Psi(\vec{r},t)}{\partial t}$$

(provided $E = \hbar\omega = \frac{p^2}{2m} = \frac{\hbar^2 k^2}{2m}$)

Stationary Schrödinger's Equation (I)

$$-\frac{\hbar^2}{2m}\nabla^2\Psi(\vec{r},t)+V(\vec{r},t)\Psi(\vec{r},t)=i\hbar\frac{\partial\Psi(\vec{r},t)}{\partial t}$$

Stationary Schrödinger's Equation (II)

$$\left[-\frac{\hbar^2}{2m}\nabla^2 + V(\vec{r})\right]\varphi(\vec{r}) = E\varphi(\vec{r})$$